

Role of solvent-ligand exchange in the reaction of iron(II) polypyridyl complexes with H₂O₂

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Iron polypyridyl complexes have been shown to be potent catalysts for epoxidation and *cis*-dihydroxylation of alkenes as well as oxidation of alkyl C-H bonds to alcohols and ketones using H₂O₂.^[1] Iron complexes with Tetradentate nitrogen ligands such as TPA and BPMEN have been used in the stereoselective epoxidation using H₂O₂.^[2] The N₃Py ligands, such as Bn-N₃Py (figure 1) have been used for epoxidation also and have shown a strong dependence on solvent in their reactivity.^[3] The complex [Fe^{II}(BnN₃Py)(CH₃CN)₂](ClO₄)₂ (**1**) provided *cis*-diol products in acetonitrile and *trans*-diols in acetone for the same substrate. Complex **1** has two labile CH₃CN ligands, one of which is easily displaced by water. In this presentation, we will discuss ligand exchange and its effect on electrochemical and spectroscopic properties and its role in allowing these complexes to react with H₂O₂.

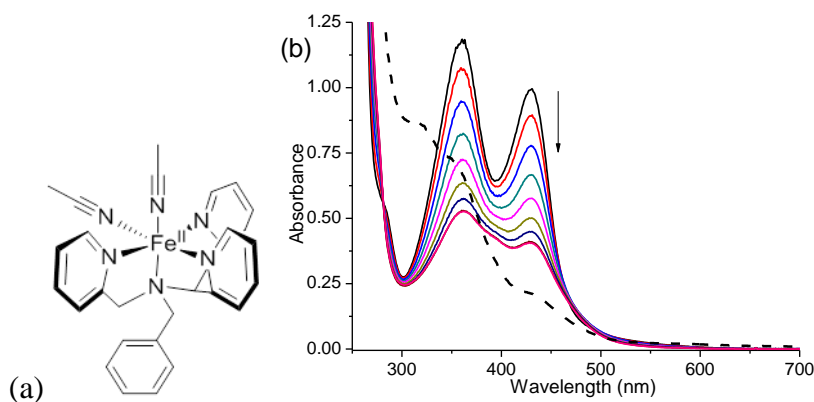


Fig 1: (a) Complex **1** ([Fe^{II}(BnN₃Py)(CH₃CN)₂](ClO₄)₂) and (b) Changes observed in the UV/Vis absorption spectrum of **1** (0.5 mM) in acetonitrile upon addition of up to 10 vol% water (solid lines) and after addition of 1 eq of H₂O₂ (dashed)

References

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